



Mark Scheme (Results)

Summer 2021

Pearson Edexcel International Advanced Level
In Chemistry (WCH16)
Paper 01: Practical Skills in Chemistry II

Question Number	Answer	Additional guidance	Mark
1(a)(i)	$\text{VO}_3^- + 2\text{H}^+ \rightarrow \text{VO}_2^+ + \text{H}_2\text{O}$	Allow multiples Ignore state symbols even if incorrect Do not award uncanceled SO_4^{2-} ions	1

Question Number	Answer	Additional guidance	Mark
1(a)(ii)	yellow	Ignore pale/light/dark/bright Do not award any other colour	1

Question Number	Answer	Additional guidance	Mark
1(a)(iii)	An explanation that makes reference to the following points: - correct colours of the oxidation states of vanadium (1) - correct explanation for first / initial green (1)	+5 (oxidation state of vanadium) is yellow and +4 is blue and +3 is green and +2 is violet Allow starting colour/answer to (a)(ii) for yellow Allow just all oxidation states/species have the correct colours Initial green is due to mixture of VO_2^+ and VO^{2+} (rather than V^{3+}) Accept initial green is due to mixture of +5 and +4 oxidation states / mixture of yellow and blue Allow vanadium cannot be oxidised from +3 to +4 in these conditions / by zinc	2

Question Number	Answer	Additional guidance	Mark
1(a)(iv)	An explanation which makes reference to the following: - oxidation of vanadium (from +2 to +3) by oxygen/O₂ (1) - oxygen/O₂ isn't a strong enough oxidising agent to oxidise vanadium(III) (under these conditions) (1)	Allow air for oxygen Allow aerial oxidation Do not award +2 to +4/+5 Standalone mark Allow oxygen/O₂ cannot oxidise +3 Allow oxidation to +4/+5 has a high activation energy Allow oxidation to +4/+5 is too slow Allow any indication that no further oxidation (of +3) occurs eg V³⁺ ions are harder to oxidise Ignore just no further reaction occurs Ignore just V³⁺ is stable	2

Question Number	Answer	Additional guidance	Mark
1(b)(i)	Any two from: [CuCl₄]²⁻ (1) [Cu(H₂O)₆]²⁺ (1) [Cu(H₂O)₅Cl]⁺ (1)	Ignore missing square brackets Do not award any complexes containing NH₃/NH₄⁺ Accept [CuCl₃]⁻ / [Cu(H₂O)₃Cl₃]⁻ Do not award [CuCl₂]⁻ Allow [Cu(H₂O)₄]²⁺ Comment allow correct names: tetrachlorocuprate(II) hexaaquacopper(II) pentaquachlorocopper(II)	2

Question Number	Answer	Additional guidance	Mark
1(b)(ii)	turns (from blue-green to) green	Accept turns green then yellow Accept turns yellow Allow turns green-yellow or any combination	1

Question Number	Answer	Additional guidance	Mark
1(b)(iii)	A description which includes: (pale) blue precipitate (of copper(II)) hydroxide)	Allow solid/ppt/ppte for precipitate Ignore gas evolved Ignore deep blue solution Do not award effervescence Do not award incorrect name/formula of precipitate	1

Question Number	Answer	Additional guidance	Mark
1(b)(iv)	An answer which makes reference to the following points: (gas evolved is) ammonia (test for ammonia) turns (damp red) litmus paper blue OR produces white smoke with HCl	(1) Accept NH_3 Allow just $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$ Allow turns universal indicator paper blue (1) Do not award white/misty fumes Ignore $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$	2

Question Number	Answer	Additional guidance	Mark
1(c)	An explanation that makes reference to the following points: • (formation of) ethanoic acid / CH_3COOH (on addition of concentrated sulfuric acid) (1) • (formation of) ester / ethyl ethanoate (on addition of ethanol) (1) • anion Y^- is CH_3COO^- / ethanoate (1)	Accept acetic acid Allow just carboxylic acid Accept $\text{CH}_3\text{COOC}_2\text{H}_5$ Accept ethyl acetate Allow name or formula of any ethyl ester Accept salt is ammonium ethanoate/$\text{CH}_3\text{CO}_2\text{NH}_4$ Accept ammonium acetate Allow name or formula of any carboxylate ion containing between one and four carbon atoms	3

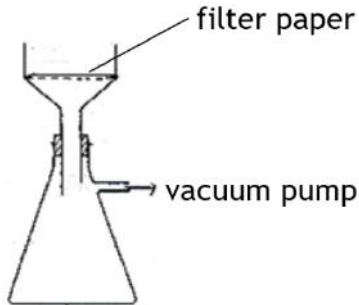
(Total for Question 1 = 15 marks)

Question Number	Answer	Additional guidance	Mark
2(a)	Any one of: - quicker - do not need the accuracy of a graduated pipette	Allow volume is approximate/to 1SF Ignore there are not 8 cm ³ pipettes Ignore water is in excess	1

Question Number	Answer	Additional guidance	Mark
2(b)	dark blue	Allow deep/royal for dark Do not award 'just' blue Do not award mention of any other colour Ignore any reference to the formation of a precipitate that subsequently dissolves	1

Question Number	Answer	Additional guidance	Mark
2(c)	ammonia/NH₃ (gas) is toxic	Accept poisonous/corrosive/irritant for toxic Ignore harmful/dangerous/health hazard	1

Question Number	Answer	Additional guidance	Mark
2(d)	the tetraamminecopper(II) sulfate-1-water/ $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ is less soluble in (aqueous) ethanol (than water so crystallises)	Allow product/salt/complex/crystals/ precipitate/solid/it for tetraamminecopper(II) sulfate-1-water Allow insoluble in ethanol	1

Question Number	Answer	Additional guidance	Mark
2(e)	A diagram showing: - Buchner funnel and labelled filter paper (1) - Buchner flask and (rubber) seal (1) (side arm with) vacuum pump (1)	Example of diagram:  Funnel must show perforations/holes below the filter paper Allow any properly shaped Buchner funnel Allow sintered glass funnel Do not award porous paper Do not award fluted filter paper Allow conical flask with side arm Allow vacuum/pump/reduced pressure/aspirator/suction Ignore just water tap Do not award pressure out/negative pressure	3

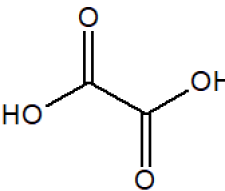
Question Number	Answer	Additional guidance	Mark
2(f)(i)	to remove (soluble) impurities	Ignore to wash the crystals	1

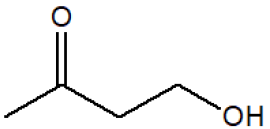
Question Number	Answer	Additional guidance	Mark
2(f)(ii)	hot ethanol would dissolve the tetraamminecopper(II) sulfate-1-water/$\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ OR so only a very small/the minimum amount of tetraamminecopper(II) sulfate-1-water/$\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ dissolves (in cold ethanol)	Allow product/salt/complex/crystals/precipitate/solid/it for tetraamminecopper(II) sulfate-1-water Allow just so it does not dissolve Allow just it is less soluble in cold ethanol Do not award insoluble in ethanol Ignore just to minimise loss of product	1

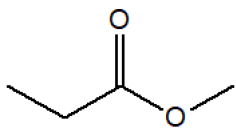
Question Number	Answer	Additional guidance	Mark
2(g)(i)	M_r values of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ (1) - mols of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and mols $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ OR theoretical mass $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ (1) - percentage yield to 2SF or 3SF (1)	Example of calculation: $M_r \text{ CuSO}_4 \cdot 5\text{H}_2\text{O} = 249.6$ Allow 249.5 $M_r \text{ Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O} = 245.6$ Allow 245.5 Mols $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ = $\frac{2.17}{249.6}$ (= $0.0086939 / 8.6939 \times 10^{-3}$) Mols $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ = $\frac{2.54}{245.6}$ (= $0.010342 / 1.0342 \times 10^{-2}$) Theoretical mass $\text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ = $0.0086939 \times 245.6 = 2.1352 \text{ (g)}$ TE on M1 % yield = $0.010342 / 0.0086939 \times 100$ = 118.96 = 119%/120% OR % yield = $2.54 / 2.1352 \times 100$ = 119%/120% Allow 119.0% TE on M2 Correct answer with some working scores (3) Just $2.54 / 2.17 \times 100 = 117\% / 120\%$ scores (0) Just $2.17 / 2.54 \times 100 = 85.4\% / 85\%$ scores (0) If no other mark awarded, M_r and mols of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} / \text{Cu}(\text{NH}_3)_4\text{SO}_4 \cdot \text{H}_2\text{O}$ scores (1)	3

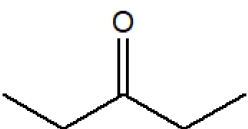
Question Number	Answer	Additional guidance	Mark
2(g)(ii)	An answer which makes reference to the following point: damp crystals	Allow wet/not properly dried/some ethanol/water remains Allow product etc for crystals Ignore just impurities Do not award it is a hydrated salt/has water of crystallisation	1

(Total for Question 2 = 13 marks)

Question Number	Answer	Additional guidance	Mark
3(a)(i)	(Compound) E	Accept correct structure: 	1

Question Number	Answer	Additional guidance	Mark
3(a)(ii)	(Compound) B	Accept correct structure: 	1

Question Number	Answer	Additional guidance	Mark
3(a)(iii)	(Compound) F	Accept correct structure: 	1

Question Number	Answer	Additional guidance	Mark
3(a)(iv)	(Compound) D	Accept correct structure: 	1

Question Number	Answer	Additional guidance	Mark
3(b)(i)	An answer that makes reference to the following points:	Result dependent on suitable test If two or more tests given, all results must be correct to score (2)	2
	- chemical test (1) - result of the selected test with A and B (1)	Examples of correct answers:	
	Chemical test	Result with A and B	
	(heat with) sodium dichromate(VI)/Na ₂ Cr ₂ O ₇ and sulfuric acid/H ₂ SO ₄ Allow just acidified dichromate / H ⁺ and Cr ₂ O ₇ ²⁻	(solution changes from orange to) green/blue with B (and no change with A)	
		OR	
	metal carbonate/metal hydrogencarbonate by name or formula	effervescence/fizzing/bubbles with A (and no change with B)	
		OR	
	magnesium/Mg	effervescence/fizzing/bubbles with A (and no change with B)	
		OR	
	ethanol/C ₂ H ₅ OH and a strong acid (by name or formula) and warm Allow just H ⁺ for strong acid	fruity smell with A (and no change with B)	
		OR	
	ethanoic acid/CH ₃ COOH and a strong acid (by name or formula) and warm Allow just H ⁺ for strong acid	fruity smell with B (and no change with A)	
		Do not award sodium Do not award PCl ₅ Do not award iodoform test Do not award Brady's reagent/2,4-DNP(H)	

Question Number	Answer	Additional guidance	Mark
3(b)(ii)	An answer that makes reference to the following points: - suitable test (1) - result of the selected test with C and D (1)	Result dependent on test (Warm with) iodine/I₂ and (aqueous) sodium hydroxide/NaOH/alkali Allow iodoform test Accept potassium iodide/KI and sodium chlorate((I))/NaClO (Pale) yellow precipitate with C (and no change with D) Allow antiseptic smell with C (and no change with D) If no other mark awarded, Brady's reagent/2,4-DNP(H) and measure melting temperature of (purified orange) solid and compare with data book scores (1)	2

Question Number	Answer	Additional guidance	Mark
3(c)(i)	(the expansion of trapped) air		1

Question Number	Answer	Additional guidance	Mark
3(c)(ii)	heat is distributed more uniformly/evenly (by convection)	Allow the temperature is more even/uniform Allow the temperature measurement is more accurate Allow the temperature rises more gradually Ignore references to evaporation	1

Question Number	Answer	Additional guidance	Mark
3(c)(iii)	the boiling temperature of compound A is higher than 100°C/water	Allow the boiling temperature of mineral oil is higher than water Allow mineral oil boils above 180°C Allow mineral oil boils at a higher temperature than compound A Allow water boils below 120°C Ignore just water boils at 100°C Ignore references to evaporation	1

Question Number	Answer	Additional guidance	Mark
3(c)(iv)	(boiling temperature depends on atmospheric) pressure (which) is variable	Allow boiling temperature is pressure dependent Ignore references to variation in just conditions/temperature	1

Question Number	Answer	Additional guidance	Mark
3(d)(i)	- solid M (anhydrous) calcium chloride/CaCl₂___ (1) - solid N soda lime_____ (1)	If name and formula given, both must be correct Allow (anhydrous) calcium sulfate/ sodium sulfate/magnesium sulfate/silica gel Do not award sulfuric acid/copper sulfate/ cobalt chloride/calcium oxide Allow potassium hydroxide/sodium hydroxide/ calcium hydroxide/calcium oxide Do not award limewater Correct substances in reverse order scores (1)	2

Question Number	Answer	Additional guidance	Mark
3(d)(ii)	- mass of hydrogen - OR moles hydrogen - mass of carbon - OR moles carbon (dioxide) - mass of oxygen - OR % mass of oxygen - calculated empirical formula	Example of calculation: mass H = $\frac{2}{18} \times 1.28 = 0.14222$ (g) OR moles H = $\frac{1.28}{18} \times 2 = 0.14222$ (mols) mass C = $\frac{12}{44} \times 3.14 = 0.85636$ (g) OR moles C/CO₂ = $\frac{3.14}{44} = 0.071364$ (mols) mass O = $1.57 - 0.14222 - 0.85636 = 0.57142/0.57$ (g) OR % mass O = $100 - 9.0587 - 54.545 = 36.396/36\%$ TE on M1 and M2 provided answer is positive $\begin{array}{ccccc} \text{C} & : & \text{H} & : & \text{O} \\ \frac{0.85636}{12} & : & \frac{0.14222}{1} & : & \frac{0.57142}{16} \\ 0.071363 & : & 0.14222 & : & 0.035714 \\ 2 & : & 4 & : & 1 \end{array}$ empirical formula is C₂H₄O Allow use of percentage masses in ratio TE on M1, M2 and M3 Ignore SF except 1SF in mass and moles	4

		Max (2) (M3 and M4) if 1.28 g and 3.14 g confused giving empirical formula CH_{12}O_2 If no other marks awarded, for 1.28 g and 3.14 g confused: mass/moles H = 0.348889 AND mass C = 0.349091 OR moles C = 0.029091 scores (1) If no other mark awarded, correct empirical formula scores (1) Comment empirical formula is $\text{C}_2\text{H}_4\text{O}$ can be awarded if seen in (d)(iv) provided mole ratio correctly calculated	
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Question Number	Answer	Additional guidance	Mark
3(d)(iii)	$(m/z =) 88$	Accept answer clearly annotated on mass spectrum	1

Question Number	Answer	Additional guidance	Mark
3(d)(iv)	molecular formula	$\left(x = \frac{M_r}{M_r(\text{C}_2\text{H}_4\text{O})} = \frac{88}{44} = 2\right)$ molecular formula is $\text{C}_4\text{H}_8\text{O}_2$ No TE on (d)(ii) or (d)(iii)	1

Question Number	Answer	Additional guidance	Mark
3(d)(v)	An answer which makes reference to the following points: • identification of compound F (1) • justification with reference to both molecular formula/M_r AND fragmentation pattern (1)	No TE on (d)(iv) Examples of justification: peak <u>s</u> at $m/z = 29$ (for $C_2H_5^+$) OR peak at $m/z = 59$ (for $COOCH_3^+$) OR no peak <u>s</u> at $m/z = 43$ (for CH_3CO^+) OR no peak at $m/z = 45$ (for $C_2H_4OH^+$) AND molecular formula $C_4H_8O_2$ / $M_r = 88$ peak <u>s</u> at $m/z = 29$ (for $C_2H_5^+$) AND D does not have molecular formula $C_4H_8O_2$ / $M_r = 88$ <u>Ignore reference to peaks at $m/z = 31/57$</u> F as has peaks at $m/z = 29$ AND $m/z = 59$ scores (2) If neither M1 nor M2 awarded, any of the following scores (1) B as has molecular formula $C_4H_8O_2$ / $M_r = 88$ OR A as has $M_r = 88$ OR D as has a peak <u>s</u> at $m/z = 29$	2

(Total for Question 3 = 22 marks)

Total for Paper = 50 marks